

Periodic Trends Yahoo Answers

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Periodic Trends Yahoo Answers

Explain each of the following comparisons. Calcium has a smaller second ionization energy than does potassium. Lithium has a larger first ionization energy than does cesium. Magnesium has a larger third ionization energy than does aluminum.

Periodic Trends? | Yahoo Answers

Describe the following Periodic Trends, and give a theoretical explanation for each trend. a) atomic radius b) ionization energy c) electro negativity d) electron affinity How do we use theoretical explanation to describe this and in general Please and Thank You

Describing Periodic Trends? | Yahoo Answers

periodic trends? can anyone explain to me? how to know the end of electronic configuration of an element without looking at the atomic number of that element, just the given group and period? ... Join Yahoo Answers and get 100 points today. Join. Trending Questions. Trending Questions.

periodic trends? | Yahoo Answers

The properties of the elements exhibit trends. These trends can be predicted using the periodic table and can be explained and understood by analysing the electron configurations of the elements. Elements tend to gain or lose valence electrons to achieve stable octet formation. Stable octets are seen in the inert gases, or noble gases, of Group VIII of the periodic table.

Trends of the periodic table? | Yahoo Answers

a) $Mg < Sr < Ba$. Radius increases every time you add an electron shell. So as you move down within a group, the radius increases. b) $K < Kr < Ar$. First ionization energy increases for the main group elements as the shell fills, so within a period elements on the left will have lower IE1 than elements to the right.

Periodic Trends (Rank these in order of ... - Yahoo Answers

Hello, this will probably all you need to know for periodic table trends: Atomic Radii (also sometimes referred to as Atomic Size): The atomic radii of elements on the periodic table increase and you go from the right of the periodic table to the left of the periodic table. Also, as you go from top of the table to the bottom, it also increases.

what are the TRENDS in the periodic table? | Yahoo Answers

The experiment we did was about Periodic Trends. We determined the trend of reactivity of nonmetals down a group, the trend of reactivity of metals along a period and the trend of acidic properties of elements along a period. The halogens used were Cl_2 , Br_2 and I_2 . Metals used were Sodium, Magnesium (magnesium ribbon), Aluminum.

Periodic Trends experiment question? | Yahoo Answers

1. Which of the following ions (Na^+ , Cl^- , Ca^{2+} , and/or S^{2+}) are isoelectronic with Ar? My Answer: Ca^{2+} , S^{2+} (WRONG) 2. Rank these elements (Kr, O, F) according to electron affinity by most energy released by gaining an electron or most energy absorbed by gaining an electron. My Answer: $Kr > O > F$ (WRONG) 3. Predict the ground-state electron configuration of the following ions. Write your ...

HELP!!! Periodic Trends? | Yahoo Answers

The human nature has been to classify the known substances on the basis of nature of the substance, for example substances are classified as solids liquids and gasses, metals non metals, pure substance which could not be further subdivide into other substances were named as elements, and in combined form as compounds, Hydrogen and oxygen are elements and water is a compound made of Hydrogen and ...

Periodic trends for atomic number and ... - Yahoo Answers

Please answer: 1) What is ionization energy? 2) What is electron affinity? 3) Which criterion for metals and non-metals is more reliable? 4) What is the trend of reducing property within a group? within a period? Thanks! :DD

Please help! Periodic trends! | Yahoo Answers

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Periodic Trends Yahoo Answers - contradatrinitas.it

how do all the trends in the periodic table connect? Answer Save. 1 Answer. ... + 100. Join Yahoo Answers and get 100 points today. Join. Trending questions. Trending questions. What is wrong with trump? 23 answers. Determine the molar mass of an ideal gas B if 0.622 g sample of gas B occupies a volume of 300 mL at 35 °C and 1.038 atm.?

how do all the trends in the periodic table connect ...

electronegativity relates to how attractive is the electron to atom of element. ionization energy- minimal energy required to remove (to infinity) an electron from the atom or molecule isolated in free space and in its ground electronic state

How do periodic trends in ... - Yahoo Answers

(A)Classify the elements according to whether they are a metal, a semimetal, or a nonmetal. Ga, Si, Te, Sn, P (B)Classify the periodic trends according to whether they increase or decrease from left to right. ionization energy, electronegativity, atomic radius, metallic character (C)Classify the periodic trends according to whether they increase or decrease from top to bottom.

Classify the elements according to whether ... - Yahoo Answers

"Family" and "periodic trends" are just as confusing. Do your future self a giant favor and ask your instructor to teach you using only the correct terminology -- namely, Groups and Periods. Elements from each Group have similar properties because they all have the same number of electrons in their outermost shell and thus available for forming chemical bonds.

Why elements from each family have similar ... - Yahoo Answers

Ionization energy is the energy needed to remove electrons from an atom, therefore, the closer an atom is to its noble gas (the further right on the periodic table), the more energy is needed to remove that electron (because of size, valence model, affinity, etc.)

Explain the ionization energy trends in ... - Yahoo Answers

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Periodic trends: Electronegativity answers. Name ____ 1. What is the electronegativity of an ... Deselect all and choose any one group in the periodic table, with the same x-and y-axes that you used above. Look at a few more groups, but avoid the transition metals (i.e., groups 3-12).

Periodic trends: Electronegativity answers. Name

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