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energy of activation for the reaction is  $190 \text{ kJ mol}^{-1}$ . Chemical Kinetics - Example : Solved Example Problems. 1. The rate law for a reaction of A, B and C has been found to be  $\text{rate} = k [A]^2 [B][L]^{3/2}$ . How would the rate of reaction change when (i) Concentration of [L] is quadrupled. Solution

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of how experimental conditions influence the speed of a chemical reaction and yield information about the reaction's mechanism and ...

## **Chemical kinetics - Wikipedia**

Some useful references. Book: An Introduction to Chemical Kinetics (Margaret Wright, 2004) - get Book



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chapter: Rates and Mechanism of Chemical Reactions (Chapter 22 of Chemical Principles, 3rd Ed (1979) by Dickerson, Gray, and Haight.) Perfectly good for first-year general chemistry courses. Online book chapter: Reaction Rates (Concept Development Studies in Chemistry, John Hutchinson)

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