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set (57) Dalton. The  
atom is a tiny,  
indestructible particle  
with no internal  
structure. Thomson.  
The atom is a sphere of  
positive electrical  
charge with electrons  
embedded in the  
sphere.

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CHAPTER 5 What You'll  
Learn You will compare

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the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons in atoms through orbital

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in Atoms. Chapter 5 -  
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Jennie L. Borders.

Section 5.1 - Models of  
the Atom. The

Rutherford's model of  
the atom did not

explain how an atom  
can emit light or the

chemical properties of  
an atom. Plum Pudding

Model Rutherford's

Model. The Bohr Model.



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Answers  
...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how:

Aufbau principle - electrons enter the lowest energy first.

This causes difficulties because of the overlap of orbitals of different energies - follow the diagram! Pauli

Exclusion Principle - at

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...

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Atoms Models of the  
Atom Rutherford used  
existing ideas about  
the atom and proposed  
an atomic model in

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which the electrons move around the nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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Cristina Chemistry I H.

Terms in this set (82)

Rutherford's model-  
also called nuclear  
model-did not explain  
how the electrons were  
arranged

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in Atoms. Chapter 5 -  
Electrons in Atoms.

Section 5.1 - Models of

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the Atom. The Rutherford's model of the atom did not explain how an atom can emit light or the chemical properties of an atom. Plum Pudding Model Rutherford's Model. The Bohr Model.

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#### Electrons in Atoms

#### Section 55.1.1 Figure

5.1 Different elements can have similar reactions with water.

**Objectives** Compare the wave and particle natures of light. Define a quantum of energy, and explain how it is related to an energy change of matter.

Contrast continuous electromagnetic



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IB Chapter Questions.  
02:25. Problem 1

Objects get their colors  
from reflecting only  
certain wavelengths  
when hit with white  
light. Light reflected  
from a green leaf is  
found to have a  
wavelength of  $4.90$   
 $\times 10^{-7}$   
 $\text{m}$ . What  
is the frequency of the

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light? ...  
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arranged in an atom?  
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Chapter 5“Electrons in  
Atoms ... Describe the  
energies and positions  
of electrons according  
to the quantum  
mechanical model.  
Section 5.1Models of  
the Atom. OBJECTIVES:  
Describe how the  
shapes of orbitals  
related to different

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sublevels differ. Ernest Rutherford's Model.

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can each p orbital  
hold? Chapter 5:  
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